

CHAPTER 02

Acids, Bases and Salts

Salts:

Salts can be described as the chemicals, which are formed by the combination of acids and bases.

Depending upon their p^H value salts can be classified as acidic salts, basic salts and neutral salts.

Acidic salts:

The salts, which are formed by the combination of strong acid and weak base, are described as acidic salts. Solutions of acidic salts will be acidic in nature means their p^H will be less than 7.

Ex.

Aluminium chloride. $AlCl_3$

Zinc sulphate. $ZnSO_4$ etc.

Basic salts:

The salts, which are formed due to combination of a weak acid and a strong base, are basic salts. The

solutions of basic salts will be basic in nature and their pH is greater than 7.

Ex.

Sodium acetate. CH_3COONa

Potassium acetate. CH_3COOK etc.

Neutral salts:

The salts, which are formed due to combination of strong acid and strong base. The solutions are neutral in nature and their p^{H} is equal to 7.

Ex.

Sodium chloride. NaCl

Potassium nitrate. KNO_3

Chemicals obtained from common salt:

Common salt is generally obtained from seawater. Seawater is the source of many different sea salts. Common salt is mined from seawater and contains large crystals, which appear in brown colour. Brown colour of common salt is due to presence of impurities.

Large crystals of salt obtained from seawater are called rock salt. Rock salt is the source of chemicals like sodium hydroxide (NaOH), Baking soda (NaHCO_3), Washing soda (Na_2CO_3), Calcium sulphate (CaSO_4) etc.

